

Bell Work

27-Nov-2017

How many valence electrons do magnesium and oxygen have?

Write out the ion for each compound($X^{-/+c}$).

What do you think their chemical formula would be if they combine?

Objective:

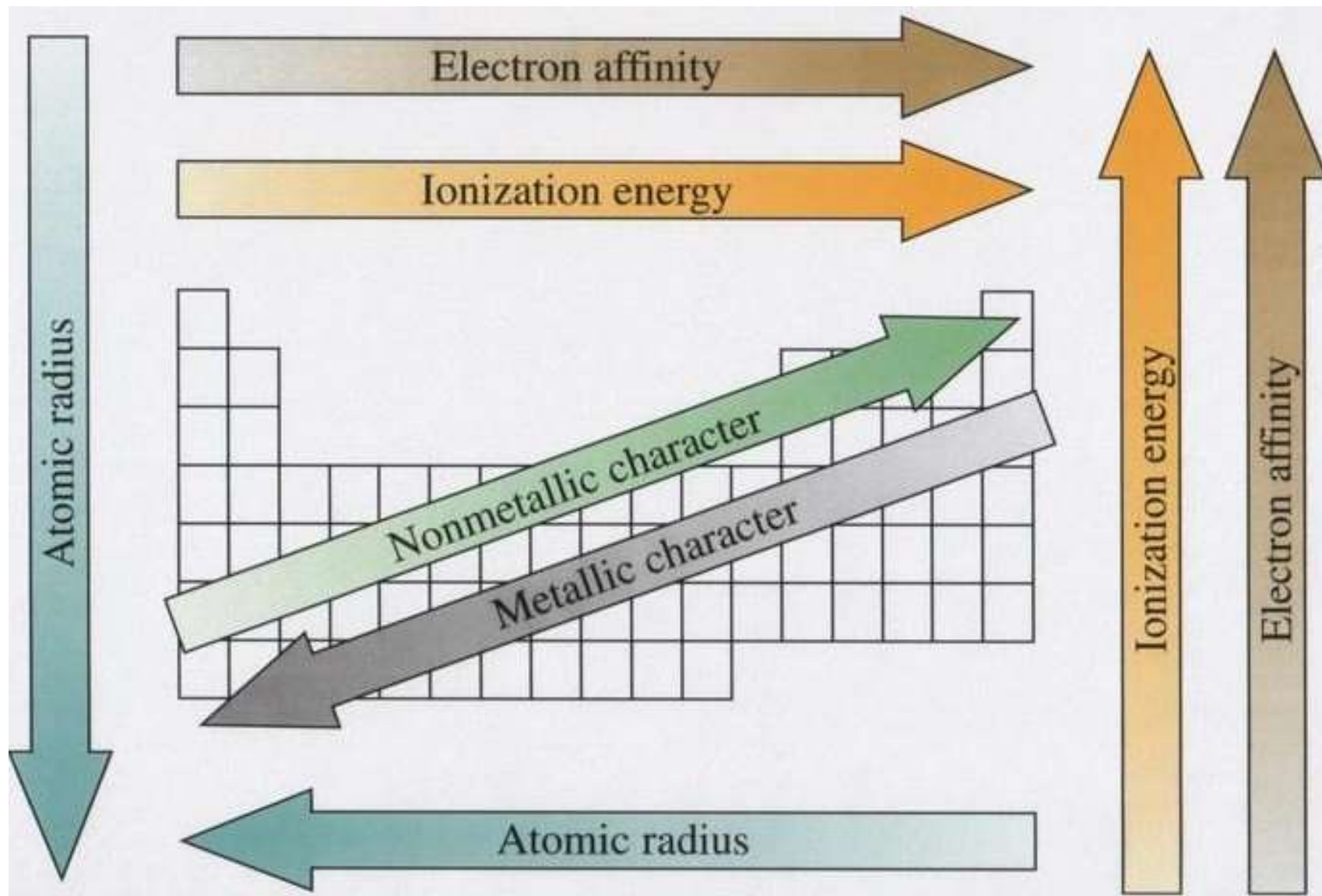
You will UNDERSTAND how to write a basic chemical formula

Ion Chips

At your lab bench go to class website and move through the notes for 27Nov17. you should understand the basics of and ionic bond. These are your note...

When finished go through and start the Ion Chip activity.

Periodic Trends



Ions

Some compounds are composed of particles called “ions”

An **ion** is an atom (or group of atoms) that has a **positive** or **negative charge**

Atoms are neutral because the number of protons equals electrons

Positive and **negative** ions are formed when **electrons** are **transferred** (lost or gained) between atoms

Ionic Compounds

Ionic compounds contain ionic bonds

Formed when e^- are given/ taken between two atoms.

e^- are exchanged between atoms so that each atom will have a full outer shell (**octet rule**).

Ionic Compounds

When e^- are given/ taken, ions are formed, & the + ions attract the - ions.

Positive ion = Cation

(Usually a metal, no change in name)

Negative ion = Anion

(Usually a nonmetal, usually ends in ide, ate, or ite)

Ionization energy/ Electronegativity Trend

Increases



The Periodic Table of the Elements

1 H Hydrogen 1.00794																	2 He Helium 4.003				
3 Li Lithium 6.941	4 Be Beryllium 9.012182															5 B Boron 10.811	6 C Carbon 12.0107	7 N Nitrogen 14.00674	8 O Oxygen 15.9994	9 F Fluorine 18.9984032	10 Ne Neon 20.1797
11 Na Sodium 22.989770	12 Mg Magnesium 24.3050															13 Al Aluminum 26.981538	14 Si Silicon 28.0855	15 P Phosphorus 30.973761	16 S Sulfur 32.066	17 Cl Chlorine 35.4527	18 Ar Argon 39.948
19 K Potassium 39.0983	20 Ca Calcium 40.078	21 Sc Scandium 44.955910	22 Ti Titanium 47.867	23 V Vanadium 50.9415	24 Cr Chromium 51.9961	25 Mn Manganese 54.938049	26 Fe Iron 55.845	27 Co Cobalt 58.933200	28 Ni Nickel 58.6934	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.61	33 As Arsenic 74.92160	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80				
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.90585	40 Zr Zirconium 91.224	41 Nb Niobium 92.90638	42 Mo Molybdenum 95.94	43 Tc Technetium (98)	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.90550	46 Pd Palladium 106.42	47 Ag Silver 107.8682	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.90447	54 Xe Xenon 131.29				
55 Cs Cesium 132.90545	56 Ba Barium 137.327	57 La Lanthanum 138.9055	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.078	79 Au Gold 196.96655	80 Hg Mercury 200.59	81 Tl Thallium 204.3833	82 Pb Lead 207.2	83 Bi Bismuth 208.98038	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)				
87 Fr Francium (223)	88 Ra Radium (226)	89 Ac Actinium (227)	104 Rf Rutherfordium (261)	105 Db Dubnium (262)	106 Sg Seaborgium (263)	107 Bh Bohrium (262)	108 Hs Hassium (265)	109 Mt Meitnerium (266)	110 (269)	111 (272)	112 (277)	113 (284)	114 (289)								
58 Ce Cerium 140.116	59 Pr Praseodymium 140.90765	60 Nd Neodymium 144.24	61 Pm Promethium (145)	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93032	68 Er Erbium 167.26	69 Tm Thulium 168.93421	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967								
90 Th Thorium 232.0381	91 Pa Protactinium 231.03588	92 U Uranium 238.0289	93 Np Neptunium (237)	94 Pu Plutonium (244)	95 Am Americium (243)	96 Cm Curium (247)	97 Bk Berkelium (247)	98 Cf Californium (251)	99 Es Einsteinium (252)	100 Fm Fermium (257)	101 Md Mendelevium (258)	102 No Nobelium (259)	103 Lr Lawrencium (262)								

Increases

Oxidation Number Trends

+1
+2

The Periodic Table of the Elements

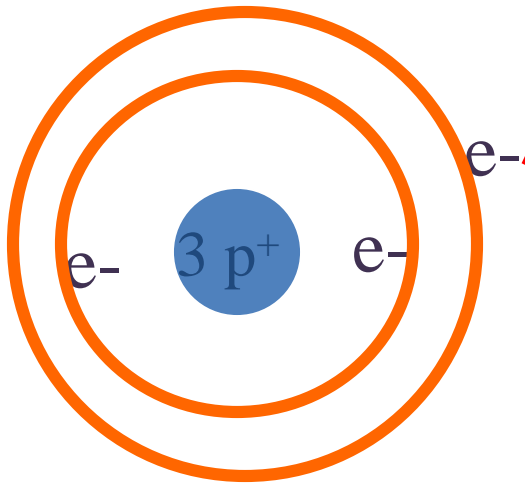
+3 **+/-** **4** **-3** **-2** **-1**

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3 Li Lithium 6.941	4 Be Beryllium 9.012182												5 B Boron 10.811	6 C Carbon 12.0107	7 N Nitrogen 14.00674	8 O Oxygen 15.9994	9 F Fluorine 18.9984032	10 Ne Neon 20.1797					
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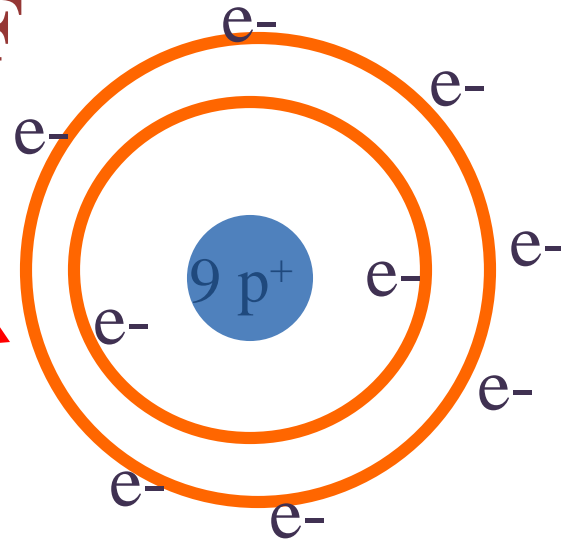
Example of how ionic bonds are formed...

Li



Wants to get rid of 1 e⁻

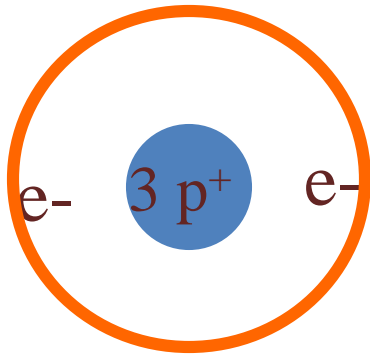
F



Wants to gain 1 e⁻

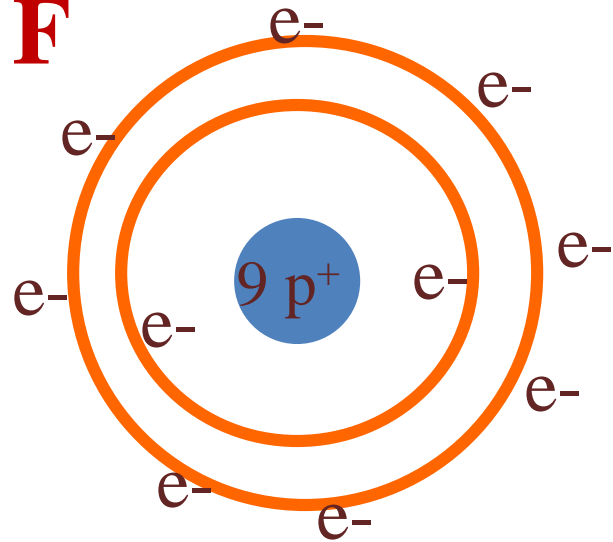
*... since there is a mutual need
to give and take one e^- ...*

Li



Gives the 1 e^- to F, to
achieve a full outer level
and...

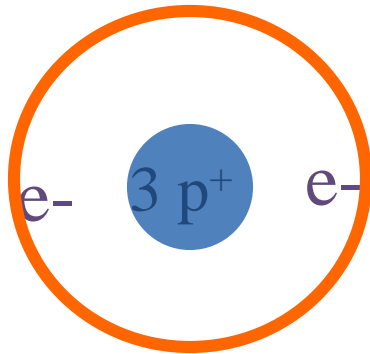
F



F gladly takes the 1
 e^- from Li to also
achieve a full outer
level.

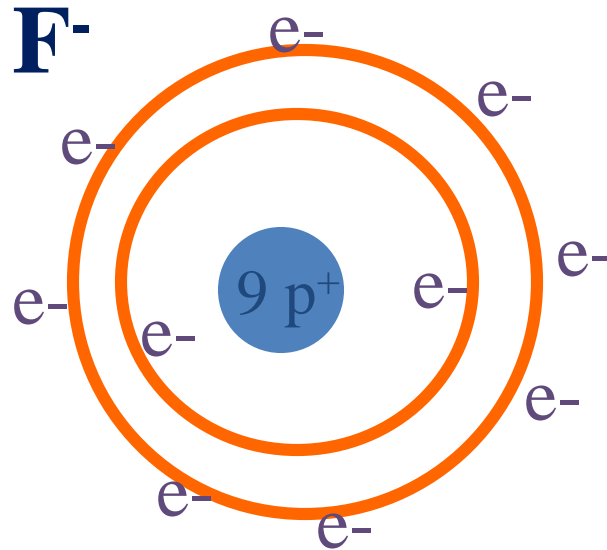
*Li and F have now become
IONS!*

Li^+



Li has 3 p+ and 2 e-

+1



F has 9 p+ and 10 e-

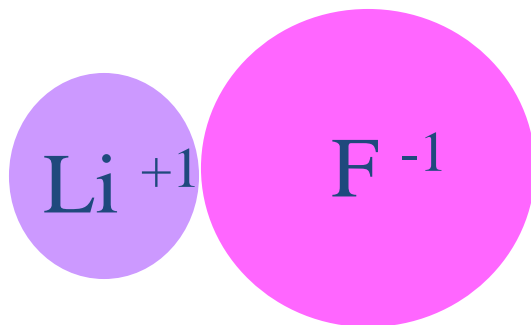
-1

Remember the rules of attraction!

Since lithium is +1 and fluorine is -1, they are attracted to each other since...

OPPOSITES ATTRACT!

Now, together, they make...



or... **LiF**

Ion Prediction

Correctly write the ion, showing oxidation number.

Ion	Formula	Cation or Anion
Nitride		
Chloride		
Sodium		
Oxide		
Calcium		
Phosphide		

Warm up

What charge do ions in the alkaline earth metals always have?

What charge do group 5, 6, and 7 element generally have?

Ion Chip Practice

Use the ion “chips” to make each of the compound on the list. Then write the correct chemical formula.

Do not loose the ions chips and return to the bag when finished.

Show all work – Ionic compound name,

Ex.